Course Code  | Course Title      | Weekly Hours* | ECTS | Weekly Class Schedule                        
-------------|------------------|---------------|------|----------------------------------------------
NS104        | General Chemistry| 3             | 2    | 6    Monday 11:00-12:50 Wednesday 9:00-9:50  
Prerequisite  |                  |               |      |                                              
Lecturer     |                  |               |      |                                              
Altijana Hromic-Jahjefendic |                |           |      |                                              
Phone        |                  |               |      |                                              
397 33 957 217 |                |           |      |                                              
E-mail       |                  |               |      |                                              
ahromic@ius.edu.ba |           |      |                                              
Course Objectives  |                  |               |      |                                              
Students will build a basic knowledge of the structure of chemistry. They will analyse scientific concepts and think critically and review the importance and relevance of chemistry in our everyday lives.  
Textbook  |                 |               |      |                                              
Learning Outcomes  | After successful completion of the course, the student will be able to: |               |      |                                              
1. Define the basic concepts of chemistry (the structure of atoms, chemical bonds, properties of elements etc.).  
2. Classify matter by its state and binding behavior using the Periodic table as a reference.  
3. Understand and explain basic equilibrium chemistry and the principle behind four main classes of reaction.  
4. Recognise in real live organic compounds and understand their functions and structures.  
5. Compare and contrast the chemical behavior and physical properties of common substances.  
Teaching Methods  | Assessment concept is based on continuous work with students during the semester. The examinations (quizzes, midterm, lab tutorials, homeworks and final exam) provide the opportunity for the students to demonstrate their understanding of the course material and their ability to apply critical thinking.  
WEEK  | TOPIC  | REFERENCE                        
Week 1  | Introduction to course content – syllabus explanation |  
Week 2  | Chemistry in Our Lives; Chemistry and Measurements | 2-23, 23-57  
Week 3  | Matter and Energy | 57-95  
Week 4  | Atoms and elements | 95-134  
Week 5  | PSE and Calculations QUIZ 1 | TBD  
Week 6  | Ionic and Molecular Compounds | TBD  
Week 7  | Chemical quantities and reactions | 211-254  
Week 8  | MIDTERM exam! |  
Week 9  | Balancing of chemical reactions | TBD  
Week 10 | Gases | 254-281  
Week 11 | Solutions | 281-322  
Week 12 | Acids and Bases/Equilibrium QUIZ 2 | 322-360  
Week 13 | Introduction to Hydrocarbons | 360-398  
Week 14 | Naming of Hydrocarbons | TBD  
Week 15 | Review | -  
Assessment Methods and Criteria  | Evaluation Tool | Quantity | Weight | Alignment with LOs |  
| Final Exam | 1 | 50 | 1,2,3,4,5 |  
| Semester Evaluation Components |  | | |  
| In-term Exam | 1 | 30 | 1,2,3,4 |  
| Lab Tutorials | 6 | 20 | 3 |  
| Quizes | 2 | 10 | 1,2,3,4 |  
**ECTS Credit Calculation**  
Activity  | Hours | Weeks | Student Workload Hours  | Activity  | Hours | Weeks | Student Workload Hours  
Lecture  | 3 | 14 | 42 | In-term exam study | 14 | 1 | 14  
Assignments  | 6 | 6 | 36 | Final exam study | 16 | 1 | 16  
In-term  | 5 | 4 | 20 | Term | 8 | 1 | 8  
Home study  | 1 | 14 | 14 |  
Total Workload Hours = 150  
ECTS Credit = 6  
Course Academic Quality Assurance: Semester Student Survey Last Update Date: 07/10/2019